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there is a brief explanation of content in physical chemistry.

Physical Chemistry Study Guide - NetSuite

$\ln K = (\quad rS/R) - (\quad rH/RT)$ The equilibrium for the dissociation of one mole of iodine vapour is 3.94×10^{-4} at

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900K and the standard enthalpy is 154 kJmol⁻¹. Calculate the equilibrium constant at 1200K assuming the standard enthalpy remains constant.

$\ln(K_2/K_1) = -\Delta_r H^\circ/R (1/T_2 - 1/T_1)$,
ANS = 6.7×10^{-2} at 1200K.

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guidance on style, word-usage conventions, nomenclature, physical quantity symbols and units, abbreviations, use of italics, and punctuation. The ACS Guide to Scholarly Communication also provides insight on what editors and reviewers look for in evaluating

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The data reported provides a guide for selecting acids to use in electrocatalysis expts. such that direct electrode redn. is avoided. ... The Journal of Physical Chemistry B 2020,

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