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NaOH: 1.009 moll 9.95 no 1.0 ml mmole of
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Reaction 1: $1 \text{ NaHCO}_3 (\text{s}) + 1 \text{ NaOH} (\text{s}) + 1 \text{ CO}_2 (\text{g})$
Reaction 2: $2 \text{ NaHCO}_3 (\text{s}) \rightarrow 1 \text{ Na}_2\text{CO}_3 (\text{s}) + 1 \text{ CO}_2 (\text{g}) + 1 \text{ H}_2\text{O} (\text{g})$
Reaction 3: $2 \text{ NaHCO}_3 (\text{s}) \rightarrow 1 \text{ Na}_2\text{O} (\text{s}) + 2 \text{ CO}_2 (\text{g}) + 1 \text{ H}_2\text{O} (\text{g})$
2. Propose a voltaic cell that would coat a piece of copper with platinum. Include the half reactions for your cell and the voltage potential based on the.

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CHEM 103 - UIUC

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Chemistry (4th Edition) Burdge, Julia
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1. Identify the group for each unknown. Your

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unknowns will be from three different groups. (Example: A K^+ salt, a halogen, and an acid.)

2. Identify the specific compound from each group. (Example: KCl , Br_2 and HCl)

3. Explain the reasoning that enabled you to come to these conclusions.

Chemistry 103 - University of Wisconsin-Eau Claire

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chemical change must be done with the best possible ...

Applied Chemistry Chemistry 101 Laboratory Manual

Lab Book: Chemistry 104 Laboratory Manual, Fall 2014 Chemistry Department, University of Wisconsin-Madison; available in the chemistry building lobby from Alpha Chi Sigma, cash only. Lab Notebook: Carbonless laboratory notebook with duplicate pages available from Alpha Chi Sigma or local bookstores.

CHEMISTRY 104

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