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(d)  $1s^2 2s^2 2p^4$  AND  $1s^2 2s^2 2p^6$  (In  
the reaction) oxygen has formed a  
negative ion (by gaining (two))

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electrons) 2 ALLOW subscripts,  
capitals ALLOW oxidation number of  
oxygen has decreased ALLOW non  
metals form negative ions IGNORE  
oxygen has gained electrons (this is  
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3(d) The electron configurations was  
frequently correctly stated, but the  
explanation for why

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